THE

BRITISH JOURNAL OF NUTRITION

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References. At the end of the paper references should be given in alphabetical order according to the name of the first author of the publication quoted, and should include the authors' initials; the title of the paper should not be included. Titles of journals should be abbreviated in accordance with the system used in the World List of Scientific Periodicals (1934: 2nd ed. Oxford University Press). Examples of such abbreviations will be found in the current numbers of the British Journal of Nutrition and a useful list has recently been published in the Journal of Physiology (1945, 104, 232). References to books and monographs should include the town of publication and the name of the publisher, as well as the date of publication and the number of the edition to which reference is made. Thus:

Barnett, J. W. & Robinson, F. A. (1942). Biochem. J. 36, 364

Culbertson, C. C. & Thomas, B. H. (1934). Rep. Ia agric. Exp. Sta. 1933-4, p. 51.

Doisy, E. A., Somogyi, M. & Shaffer, P. A. (1923). J. biol. Chem. 55, xxxi.

Fairley, N. H. (1938). *Nature, Lond.*, **142**, 1156.

Hennessy, D. J. (1941). *Industr. Engag Chem.* (Anal.)

Hennessy, D. J. (1941). Industr. Engng Chem. (Anal. ed.), 13, 216.

King, H. (1941). J. chem. Soc. p. 338.

Osborne, T. B. & Mendel, L. B. (1914a). J. biol. Chem. 17,

Osborne, T. B. & Mendel, L. B. (1914b). J. biol. Chem. 18, 1.
Osborne, T. B. & Mendel, L. B. (1916). Biochem. J. 10, 534.
Osborne, T. B., Mendel, L. B. & Ferry, E. L. (1919). J. biol. Chem. 37, 233.

Starling, E. H. (1915). Principles of Human Physiology, 2nd ed. London: Churchill.

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The methods of analysis followed should be indicated, but statistical details, such as an analysis of variance tables, need not be given unless they are relevant to the discussion. A statement that the difference between the mean values of two groups of data is statistically significant should be accompanied by an indication of the level of significance attained.

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Description of Solutions. Solutions of common acids, bases and salts are preferably defined in terms of normality (N) or molarity (M), e.g. N-HCl; 0·1 M-NaH₂PO₄. The term '%' must be used in its correct sense, i.e. g./100 g. of solution. 10% HCl means 10 g. of hydrogen chloride in 100 g. of aqueous solution, and should never be used to indicate a tenfold dilution of laboratory concentrated hydrochloric acid. For 'per cent by volume', i.e. ml./100 ml., the term '% (v/v)' may be employed. To indicate that a given weight of substance is contained in 100 ml. of solution, the term '% (w/v)' (weight per volume) may be used.

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